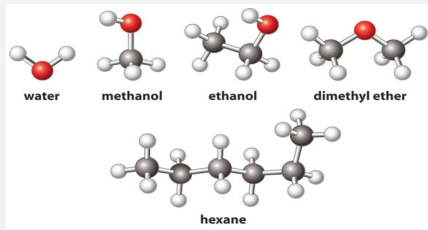


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[TP] Identify the compound with the **smallest** energy of vaporization (kJ/mol).

25% 1. water  
25% 2. ethanol  
25% 3. dimethyl ether  
25% 4. hexane



water methanol ethanol dimethyl ether  
hexane

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Response Counter 10 1

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Wednesday, October 17, 2018

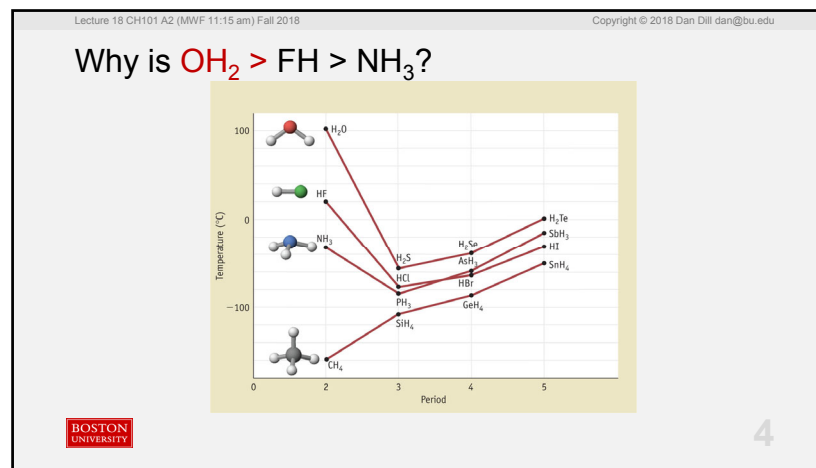
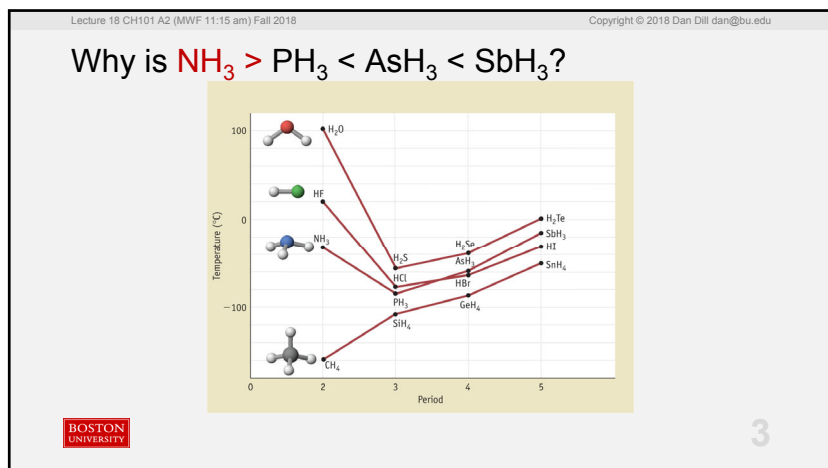
For today ...

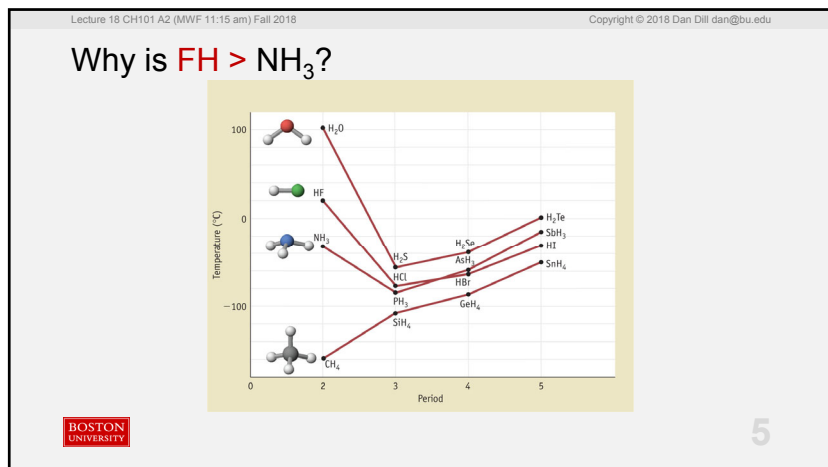
- Complete: Putting it all together: Relative boiling points
- Practice: Intermolecular forces

Next lecture: Dissolving ionic solids; Solubility guidelines fig 6.28, p 181 (memorize); Precipitation reactions; concentrations after precipitation

Suggested problems: Chapter 6: 27, 28, 31, 32, 57, 61, 62, 65, 69, 71, 73, 75, 81, 83, 85, 91, 96, 97, 99

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[TP] Identify the compound with the **smallest** energy of vaporization (kJ/mol).

25% 1. water  
25% 2. ethanol  
25% 3. dimethyl ether  
25% 4. hexane

water methanol ethanol dimethyl ether

hexane

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[TP] Identify the compound with the **largest** energy of vaporization (kJ/mol).

33% 1. Ethane,  $\text{CH}_3\text{CH}_3$   
33% 2. Propane,  $\text{CH}_3\text{CH}_2\text{CH}_3$   
33% 3. Dimethyl ether,  $\text{CH}_3\text{OCH}_3$

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[Group Quiz] The correct order of **boiling point** of  $\text{CH}_3\text{OH}$ ,  $\text{CH}_4$ ,  $\text{S}=\text{C}=\text{O}$ , and  $\text{Xe}$  is (**lowest to highest**) ...

17% 1.  $\text{CH}_3\text{OH} < \text{CH}_4 < \text{S}=\text{C}=\text{O} < \text{Xe}$   
17% 2.  $\text{Xe} < \text{S}=\text{C}=\text{O} < \text{CH}_4 < \text{CH}_3\text{OH}$   
17% 3.  $\text{Xe} < \text{CH}_4 < \text{S}=\text{C}=\text{O} < \text{CH}_3\text{OH}$   
17% 4.  $\text{CH}_4 < \text{Xe} < \text{CH}_3\text{OH} < \text{S}=\text{C}=\text{O}$   
17% 5.  $\text{CH}_4 < \text{Xe} < \text{S}=\text{C}=\text{O} < \text{CH}_3\text{OH}$   
17% 6. some other order

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