

CH 102 Quiz 9 (10 points, 10 minutes)

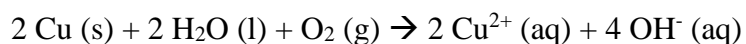
Name _____ Grading Key _____

Section/ Day _____

Cu(s) is oxidized to Cu²⁺ (aq) by oxygen gas in the basic solution.

[E_{red} (Cu²⁺/Cu) = 0.34V, E_{red} (O₂/OH⁻) = 0.40V].

1. (3points) Write the balanced net chemical reaction:



2. (3 points) Calculate E^o_{cell}.

$$E^{\circ}_{\text{cell}} = 0.40 - 0.34 = 0.06\text{V}$$

3. (4 point) Write the line notation for this reaction. Be sure to include phases.

