1. What is the expression for the equilibrium constant of an aqueous solution of the oxoacid conjugate base :OA-?
   - A $K_b = (:OA^-)(:OH^-)$
   - B $K_b = (:OH^-)(:OA^-)/(H-OA)$
   - C $K_b =(H-OA)(:OH^-)/(:OA^-)$
   - D $K_b = (H3O+)(:OA^-)/(HOA)$

2. H-OA is a weak acid with $K_a = 10^{-5}$. What is the pH of 0.1 M conjugate base :OA-?
   - A < 7
   - B 7
   - C > 7
   - D More information needed

3. H-OA is a weak acid with $K_a = 10^{-5}$ at 25°C. What is $K$ for H-OH + :OA- $\rightleftharpoons$ H-OA + :OH- at 25°C?
   - A $K > 10^{-5}$
   - B $K = 10^{-5}$
   - C $K < 10^{-5}$
   - D More information needed

4. 0.001 mol of a weak oxoacid salt, Na-OA, is placed in 1.00 L water at 25°C. Before equilibrium is established, what is [OA-]?
   - A 0
   - B 0.001
   - C 10^{-7}
   - D More information needed

5. 0.001 mol of a weak oxoacid salt, Na-OA, is placed in 1.00 L water at 25°C. Before equilibrium is established, what is [HOA]?
   - A 0
   - B 0.001
   - C 10^{-7}
   - D More information needed

6. 0.001 mol of a weak oxoacid salt, Na-OA, is placed in 1.00 L water at 25°C. Before equilibrium is established, what is [OH-]?
   - A 0
   - B 0.001
   - C 10^{-7}
   - D More information needed
7. 0.001 mol of a weak oxoacid salt, Na-OA, is placed in 1.00 L water at 25°C. After equilibrium is established, what is (HOA-)?
   A. 0
   B. ~0.001
   C. 10^-7
   D. More information needed

8. 0.001 mol of a weak oxoacid salt, Na-OA, is placed in 1.00 L water at 25°C. After equilibrium is established, what is (OA)?
   A. 0
   B. 0.001
   C. 10^-7
   D. More information needed

9. 0.001 mol of a weak oxoacid salt, Na-OA, is placed in 1.00 L water at 25°C. After equilibrium is established, what is (OH-)?
   A. 0
   B. 0.001
   C. 10^-7
   D. More information needed