

- Relationship between energy change,  $\Delta E$  (black path), and enthalpy change,  $\Delta H$  (red path), according to
- (1) whether products have more (upper figures) or less (lower figures) energy than reactants, and
  - (2) whether work (cyan path) is done on the system (left figures) or on the surrounding (right figures).
- When no work is done at constant  $P$ , then  $\Delta E$  and  $\Delta H$  have the same value.

