Lecture 4 prep: Enthalpy change of solution

See section 12.3.

Without looking at data, answer the following about LiCl(s) and KCl(s).

Explain which one should have the larger lattice enthalpy.

Explain which one should have the larger enthalpy of aquation.

Based on your explanations, can you predict whether the enthalpy of solution of LiCl(s) is exothermic?

Based on your explanations, can you predict which one should have the enthalpy of solution of larger magnitude?